

AL HUDA SCIENCE ACADEMY

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Most Important MCQs of Chemistry

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1. What this world is made of?

A. Air | B. Matter | C. Energy | D. Matter & energy

Answer: D

2. Number of states of matter that we observe in everyday life.

A. 2 | B. 3 | C. 4 | D. 5

Answer: C

3. Which branch investigate the behaviour of substance at atomic or molecular level?

A. Organic Chemistry | B. Biochemistry | C. Analytical Chemistry | D. Physical Chemistry

Answer: D

4. The study of covalent compounds of carbon, hydrogen and their derivatives deals:

A. Organic Chemistry | B. Inorganic Chemistry | C. Analytical Chemistry | D. Industrial Chemistry

Answer: A

5. Identify the naturally occurring polymer.

A. Polyvinyl chloride (PVC) | B. Protein | C. H₂O | D. None of these

Answer: B

6. Hazardous effects of shopping bags are studied in:

A. Geochemistry | B. Inorganic Chemistry | C. Analytical Chemistry | D. Environmental Chemistry

Answer: D

7. The man made polymer is:

A. Starch | B. Polystyrene | C. Protein | D. Cellulose

Answer: B

8. Simplest form of matter is:

A. Molecules | B. Mixture | C. Elements | D. Compounds

Answer: C

9. Types of mixture are:

A. 2 | B. 3 | C. 4 | D. 5

Answer: A

10. Allotropic forms of oxygen are:

A. One | B. Two | C. Three | D. Four

Answer: C

11. Which one of the following is NOT a homogeneous mixture?

A. Solution of salt | B. Solution of sugar | C. Milk | D. Rock

Answer: D

12. Identify the nature of chocolate.

A. Element | B. Compound | C. Mixture | D. Plasma

Answer: C

13. Which one of the following is NOT a crystalline form of carbon?

A. Diamond | B. Buckminster fullerene | C. Graphene | D. Coal

Answer: D

14. What is true about element?

A. It is pure substance | B. It is the simplest form of matter | C. It contain same kind of atoms | D. All of these

Answer: D

15. The symbol of sodium element is.

A. N | B. NA | C. Na | D. Ni

Answer: C

16. Which element exists in the form of monoatomic molecule?

A. Oxygen | B. Nitrogen | C. Chlorine | D. Noble gases

Answer: D

17. Identify incorrect statement about compound.

A. A compound is an impure substance | B. It is formed by chemical combination | C. Its components cannot separate by physical mean | D. Its components lose their properties

Answer: A

18. Which of the following is an impure substance:

A. Gold | B. Water | C. Ammonia | D. Air

Answer: D

19. A sample that is formed by simple mixing of their components without fixed ratio is:

A. Element | B. Compound | C. Mixture | D. Table salt

Answer: C

20. A solution in which solute particles are invisible is called:

A. False solution | B. True solution | C. Colloidal solution | D. Suspension

Answer: B

21. What statement about suspension is incorrect?

A. It is a mixture | B. Solute particles do not dissolve in solvent | C. We can see the solute particles | D. It is true solution

Answer: D

22. A mixture of chalk in water is an example of:

A. True solution | B. Colloidal solution | C. Suspension | D. Homogeneous mixture

Answer: C

23. Starch solution is:

A. True solution | B. Colloidal solution | C. Suspension | D. None of these

Answer: B

24. The solution which can dissolve more amount of solute called:

A. Saturated solution | B. Unsaturated solution | C. Super saturated solution | D. Suspension

Answer: B

25. When the tiny visible particles of a substance are dispersed through a medium, the mixture is named as:

A. True solution | B. Colloid | C. Suspension | D. Saturated solution

Answer: C

26. Colour of element sulphur is:

A. Yellow | B. White | C. Brown | D. Black

Answer: A

27. John Dalton put forward his atomic theory in:

A. 1803 | B. 1903 | C. 2005 | D. 1673

Answer: A

28. How many times a proton is heavier than an electron?

A. 1936 | B. 1836 | C. 1736 | D. 1673

Answer: B

29. When neutron was discovered?

A. 1912 | B. 1923 | C. 1933 | D. 1943

Answer: C

30. How many protons are in ${}^4_2\text{He}$ atom?

A. 2 | B. 4 | C. 4 | D. 5

Answer: A

31. How much charge carries a proton?

A. $+1.6022 \times 10^{-19}\text{C}$ | B. $-1.6022 \times 10^{-19}\text{C}$ | C. $+2.6022 \times 10^{-19}\text{C}$ | D. $-2.622 \times 10^{-19}\text{C}$

Answer: A

32. How much charge carries a neutron particle?

A. -1.6022×10^{-19} | B. $+1.6022 \times 10^{-19}$ | C. 0.0 | D. None of these

Answer: C

33. Who proposed the concept of orbits?

A. Neutron | B. Dalton | C. Bohr | D. Goldstein

Answer: C

34. Fourth shell consist of subshell:

A. s | B. s,p | C. s,p,d | D. s,p,d,f

Answer: D

35. The formula used to calculate number of electrons in each shell is:

A. n^2 | B. $2n^2$ | C. $2n^3$ | D. $4n^3$

Answer: B

36. What information was obtained from discharge tube experiments?

A. Structure of atom was discovered | B. Neutrons and protons were discovered | C. Electrons and protons were discovered | D. Presence of nucleus in an atom was discovered

Answer: C

37. What does keep the particles present in the nucleus intact?

A. Particles are held together by strong nuclear force | B. Particles are held together by weak nuclear force | C. Particles are held together by electrostatic force | D. Particles are held together by dipolar force

Answer: A

38. How do electrons keep themselves away from the oppositely charged nucleus?

A. By keeping themselves stationary | B. By revolving around the nucleus | C. Due to their wave-like nature | D. A magnetic field around the nucleus keeps them away

Answer: B

39. Fundamental particles of all types of matter are:

A. Protons | B. Neutrons | C. Electrons | D. All of these

Answer: D

40. As whole an atom carries charge:

A. Positive | B. Negative | C. Atom is neutral particle | D. None of these

Answer: C

41. Atomic number of an element is represented by:

A. A | B. Z | C. W | D. X

Answer: B

42. Number of protons in an atom is called:

A. Atomic Weight | B. Atomic Mass | C. Atomic Number | D. Atomic Radius

Answer: C

43. Sum of protons and Neutrons in an atom is called:

A. Atomic Number | B. Mass Number | C. Shell Number | D. None of these

Answer: B

44. Mass number is represented by:

A. Z | B. A | C. B | D. m

Answer: B

45. An atom has mass number 16 and atomic number is 8. How many neutrons are in it?

A. 8 | B. 18 | C. 10 | D. 16

Answer: A

46. Formula for calculations of neutrons in an atom is:

A. $N = Z - A$ | B. $N = A - Z$ | C. $Z = N + A$ | D. $A = Z - N$

Answer: B

47. The mass of an atom of an element relative to the mass of light isotope of carbon taken as 12 is called:

A. Relative atomic Mass | B. Molecular Mass | C. Consecutive Mass | D. Accurate Mass

Answer: A

48. 1 amu is equal to:

A. $1.67377 \times 10^{-27} \text{Kg}$ | B. $6.02 \times 10^{-24} \text{Kg}$ | C. $1.66 \times 10^{23} \text{g}$ | D. 1.008 g

Answer: A

49. Atomic mass of sulphur is.

A. 31.972 amu | B. 31.92 g | C. 31.92 Kg | D. All of these

Answer: A

50. What will be the relative atomic mass of nitrogen given the abundances of its two isotopes, ^{14}N and ^{15}N are 99.64 and 0.35 respectively?

A. 14.0210 | B. 14.0021 | C. 14.2100 | D. 14.1200

Answer: A

51. How many elements are in the valence shell of noble gases?

A. 4 | B. 6 | C. 7 | D. 8

Answer: D

52. The ability of an atom to get two electrons in valence shell is called:

A. Bonding rule | B. Octet rule | C. Duplet rule | D. Antibonding rule

Answer: C

53. Which element is capable of forming all the three types of bonds; covalent, coordinate covalent and ionic?

A. Carbon | B. Oxygen | C. Magnesium | D. Silicon

Answer: B

54. Which of the following bonds is expected to be the weakest?

A. C-C | B. I-I | C. O-O | D. F-F

Answer: D

55. Electronic configuration of chlorine (Cl) is:

A. 2, 8, 8 | B. 2, 8, 7 | C. 2, 8, 2 | D. 2, 8, 18

Answer: B

56. The chemical bond which is formed by mutual sharing of electrons is called:

A. Ionic bond | B. Covalent bond | C. Dative bond | D. Intermolecular forces

Answer: B

57. Identify the compound that has double covalent bond.

A. H_2O | B. C_2H_4 | C. C_2H_6 | D. C_2H_2

Answer: B

58. Which of the following has polar covalent bond?

A. CO_2 | B. H_2O | C. CH_4 | D. O_2

Answer: B

59. Example of single covalent bond is:

A. H_2O | B. Cl_2 | C. NH_3 | D. All of these

Answer: D

60. Identify the property of metals:

A. Shows good metallic luster | B. Have high melting and boiling point | C. Good conductors of Heat | D. All of these

Answer: D

61. What is the most common oxidation state of alkali metals?

A. +1 | B. +2 | C. +3 | D. +4

Answer: A

62. What is the nature of bonding in metals?

A. Ionic bonding | B. Covalent bonding | C. Metallic bonding | D. Vander waals forces

Answer: C

63. Metals are good conductor of heat and electricity due to:

A. Free electrons | B. Brittleness | C. Ductility | D. Malleability

Answer: A

64. Which metal has the lowest melting point?

A. Li | B. Na | C. K | D. Rb

Answer: D

65. Which among the following has a double covalent bond?

A. Ethane | B. Methane | C. Ethylene | D. Acetylene

Answer: C

66. Metals have tendency to:

A. To lose electrons | B. To gain electrons | C. Share electrons | D. None of these

Answer: A

67. Which one of the following is more electropositive metal?

A. Li | B. Na | C. Ca | D. Mg

Answer: B

68. Identify the gas that is produced when aluminium reacts with mineral acids.

A. O_2 | B. CO_2 | C. N_2 | D. H_2

Answer: D

69. Metals forms ion:

A. Cation | B. Anion | C. Both A and B | D. None of these

Answer: A

70. Identify the most electronegative element of the periodic table.

A. F | B. Cl | C. O | D. N

Answer: A

71. Which one of the following has weakest force of attraction?

A. Covalent bond | B. Metallic bond | C. Ionic bond | D. Hydrogen bond

Answer: D

72. Melting and boiling points of a substance depends on:

A. Ionic bond | B. Covalent bond | C. Metallic bond | D. Inter-molecular forces

Answer: D

73. The attractive force between the molecules of water is called:

A. Dipole-dipole forces | B. Hydrogen bonding | C. Instantaneous dipole - induce | D. None of these

Answer: B

74. Which one of the following contains Dipole - dipole forces?

A. H_2 | B. O_2 | C. N_2 | D. HCl

Answer: D

75. Why is H_2O a liquid while H_2S is a gas?

A. water is polar compound | B. H_2O molecule is easily freeze | C. atomic size of oxygen is smaller | D.

forces of attraction between H_2O molecules are stronger than H_2S

Answer: D

76. O_3 is the chemical formula of:

A. Oxygen molecule | B. Ozone | C. Oxide | D. Oxytocin

Answer: B

77. Ionic compounds are represented by:

A. Molecular formula | B. Symbols | C. Formula unit | D. None of these

Answer: C

78. Chemical formula of methane natural gas is:

A. CH_4 | B. C_2H_6 | C. C_2H_2 | D. C_2H_4

Answer: A

79. Which one of the following is NOT covalent compound?

A. HCl | B. HF | C. PH_3 | D. CaCl_2

Answer: D

80. Empirical formula of calcium fluoride is:

A. CaF_2 | B. Ca_2F_2 | C. CaF_3 | D. CaF

Answer: A

81. Which of the following is indicated by empirical formula?

A. Actual number of compound | B. The simplest ratio between atoms | C. The minimum ratio between the ions | D. None of these

Answer: B

82. The molecular formula of benzene is:

A. CH | B. C_6H_6 | C. C_3H_3 | D. C_5H_5

Answer: B

83. The formula of Aluminum nitride is:

A. AlO | B. Al_2N_3 | C. Al_3N_2 | D. Al_4N_2

Answer: B

84. Calcium carries charge:

A. $2+$ | B. $2-$ | C. $1-$ | D. $1+$

Answer: A

85. Which is the correct formula of calcium phosphide?

A. CaP | B. CaP_2 | C. CaP_3 | D. Ca_3P_2

Answer: D

86. The molecular formula of glucose is:

A. CH_2 | B. $\text{C}_6\text{H}_{12}\text{O}_6$ | C. CHO_2 | D. CHO

Answer: B

87. Empirical formula of benzene is CH and its molecular formula is:

A. C_2H_2 | B. C_3H_3 | C. C_6H_6 | D. C_8H_8

Answer: C

88. Molecule and empirical formula of sand is:

A. SiO_2 | B. Si_2O_2 | C. Si_2O_4 | D. SiO

Answer: A

89. $\text{Ca}(\text{SO}_4)_2$ is the formula of calcium sulphate. How many sulphur atoms are in it?

A. 1 | B. 2 | C. 8 | D. 4

Answer: B

90. How many total atom are in $\text{Ca}(\text{OH})_2$?

A. 2 | B. 3 | C. 4 | D. 5

Answer: D

91. Structural formula of 2-hexene is... What will be its empirical formula?

A. C_2H_2 | B. CH | C. C_6H_{12} | D. CH_2

Answer: D

92. 1 mole of water has number of water molecules:

A. 24 g | B. 1.66×10^{22} | C. 6.022×10^{23} | D. 18 g

Answer: C

93. Avogadro's number was introduced by:

A. Italian scientist | B. American scientist | C. German scientist | D. Pakistani scientist

Answer: A

94. How many atoms are present in one gram of H_2O ?

A. 1.002×10^{23} atoms | B. 6.022×10^{23} atoms | C. 0.334×10^{23} atoms | D. 2.004×10^{23} atoms

Answer: A

95. The quantity of a substance containing Avogadro's number of particles (N_A) is called:

A. Mole | B. Avogadro's number | C. Molecular formula | D. All of these

Answer: A

96. The mass of one mole of a substance is called:

A. Molar mass | B. Formula mass | C. Molecular mass | D. None of these

Answer: A

97. A mole of H_2SO_4 contains molecules:

A. 6.02×10^{-23} | B. 6.02×10^{23} | C. 6.022×10^{-24} | D. 6.022×10^{24}

Answer: B

98. The substances which react to give a chemical change are called:

A. Products | B. Reactants | C. Ingredients | D. None of these

Answer: B

99. One mole of calcium carbonate is equal to:

A. 100 g | B. 10 g | C. 110 g | D. 80 g

Answer: A

100. Molar mass of $C_6H_{12}O_6$ is:

A. 180 | B. 98 | C. 170 | D. 160

Answer: A

101. In which form the energy is stored in a molecule?

A. Chemical energy | B. Heat energy | C. Elastic potential | D. Mechanical energy

Answer: A

102. Which of the following best defines a system in thermodynamics?

A. The entire Universe | B. The part of Universe being studied | C. The container holding a reaction | D. The laboratory where experiments are performed

Answer: B

103. The energy which is released when bond is formed and absorbed when it is broken is called:

A. Chemical energy | B. Heat energy | C. Internal energy | D. Kinetic energy

Answer: A

104. A reaction in which heat is absorbed is called:

A. Exothermic reaction | B. Endothermic reaction | C. Burning | D. Decomposition reaction

Answer: B

105. The total heat content of a body is called:

A. Enthalpy | B. System | C. Surrounding | D. All of these

Answer: A

106. Under which conditions the standard enthalpy is measured?

A. 0°C and 1 atm | B. 25°C and 1 atm | C. 0°C and 25 atm | D. 1°C and 25 atm

Answer: B

107. Which of the following is a unit of enthalpy?

A. J/s | B. mol/L | C. KJ/mol | D. L/mol

Answer: C

108. Which theory explains how reaction occurs based on collisions between particles?

A. Atomic theory | B. Collision theory | C. Molecular theory | D. Kinetic theory

Answer: B

109. Which of the following factors does not effects the rate of chemical reactions?

A. Temperature | B. Catalyst | C. Surface area | D. Composition of container

Answer: D

110. A catalyst does what to the activation energy of reactions?

A. Increases it | B. Decreases it | C. Has no effect | D. None of these

Answer: B

111. In a chemical reaction the substance that is used up is called:

A. Product | B. Catalyst | C. Limiting reactant | D. Solvent

Answer: C

112. During glycolysis which molecule breaks?

A. Alcohol | B. Protein | C. Glucose | D. Lactic acid

Answer: C

113. During glycolysis one molecule of glucose splits into:

A. 2 molecule of pyruvate | B. CO₂ and oxygen | C. Oxygen and water | D. None of these

Answer: A

114. The primary storage form of glucose is:

A. Glycogen | B. Glycoside | C. Nucleic acids | D. All of these

Answer: A

115. In reversible reaction what does the double arrow ($\rightleftharpoons$) indicates?

A. The reaction is exothermic | B. The reaction is endothermic | C. Reaction can proceed in both directions | D. The reaction is catalyzed

Answer: C

116. Formation of ammonia in closed system is an example of:

A. Endothermic process | B. Reversible reaction | C. Irreversible reaction | D. All of these

Answer: B

117. Chemical formula of copper sulphate pentahydrate is:

A. CuSO₄.5H₂O | B. CuSO₄ | C. CuSO₄.2H₂O | D. 2H₂O

Answer: A

118. Colour of copper sulphate pentahydrate is:

A. Red | B. White | C. Blue | D. Yellow

Answer: C

119. Colour of anhydrous CuSO₄ is:

A. Blue | B. White | C. Red | D. Yellow

Answer: B

120. The colour of NO₂ gas is:

A. Brown | B. Black | C. White | D. Green

Answer: A

121. When a reaction will become a reversible one?

A. If the activation energy of the forward reaction is comparable to that of backward reaction | B. If the

activation energy... is higher than... backward... | C. If the activation energy... is lower than... backward...
| D. If the enthalpy change... is zero

Answer: A

122. The reaction in which products recombine to form reactants are called:

A. Irreversible reaction | B. Reversible reaction | C. Forward reaction | D. None of these

Answer: B

123. What happens with following reaction when pressure is increased? $N_2 + 3H_2 \rightleftharpoons 2NH_3$

A. Move in forward direction | B. Move in backward direction | C. No change | D. None of these

Answer: A

124. What is true about reversible reaction?

A. It never goes to completion | B. Burning of coal is reversible reaction | C. Rate of reaction does not effected by temperature | D. None of these

Answer: A

125. Which of the following is brown?

A. N_2O_4 | B. NO_2 | C. NH_3 | D. N_2

Answer: B

126. The reversible reaction is represented by:

A. $\rightarrow$ | B. $=$ | C. $\rightarrow$ | D. $\rightleftharpoons$

Answer: D

127. Dynamic equilibrium establishes when:

A. Rate of forward is greater than reverse... | B. Rate of forward and reverse reactions become equal | C. System is opened | D. All of these

Answer: B

128. A closed system is essential for establishing:

A. Static equilibrium | B. Dynamic equilibrium | C. No equilibrium | D. Rapid reaction

Answer: B

129. An equilibrium is reached when:

A. All reactants are consumed | B. All products are consumed | C. The rate of forward and rate of reverse reaction become equal | D. None of these

Answer: C

130. Identify the oxide that is not basic in nature?

A. CaO | B. KOH | C. MgO | D. SO_3

Answer: D

131. A natural source of tartaric acid is:

A. Amla | B. Tamarind | C. Guava | D. Orange

Answer: B

132. Which acid is not used as a food or mixed with food?

A. Tartaric acid | B. Ascorbic acid | C. Citric acid | D. Formic acid

Answer: D

133. According to Arrhenius acid is a substance which dissociates in water and give:

A. H^+ | B. H_3O | C. Cl | D. O_3

Answer: A

134. According to Arrhenius base dissociate in water to give.

A. H^+ | B. OH^- | C. A and b both | D. None of these

Answer: B

135. Example of arrhenius acid is:

A. HNO_3 | B. H_2SO_4 | C. HCN | D. All of these

Answer: D

136. Examples of weak bases are:

A. Potassium Hydroxide | B. Ammonium hydroxide | C. Aluminum hydroxide | D. Both B and C

Answer: D

137. According to Bronsted concept acid.

A. Donate proton (H^+) | B. Accept Proton | C. Both a and b | D. None of these

Answer: A

138. The gas produced by reaction of acid and base is.

A. H_2 | B. Cl_2 | C. N_2 | D. O_2

Answer: A

139. The formula of sodium carbonate is:

A. NaOH | B. $NaHCO_3$ | C. Na_2CO_3 | D. $NaOHCO_3$

Answer: C

140. The formula of potassium carbonate is:

A. $KHCO_3$ | B. $KOHCO_3$ | C. KOH | D. K_2CO_3

Answer: D

141. Salt and water produce when acid reacts with.

A. Base | B. Water | C. Glucose | D. Protein

Answer: A

142. Which one of the following base is insoluble in water?

A. $Ca(OH)_2$ | B. NaOH | C. KOH | D. $Ca(OH)_2$

Answer: D

143. What do bases react with to form ammonia gas?

A. Water | B. Acids | C. Ammonium | D. Mineral acid

Answer: C

144. Which one of the following is the base of modern periodic table?

A. Atomic number | B. Ionization energy | C. Atomic Mass | D. Atomic radius

Answer: A

145. Horizontal row of the periodic table is called:

A. Group | B. Period | C. Column | D. None of these

Answer: B

146. In which period and group will you place the element which is an important part of the solar cell?

A. Third period and 14th group | B. Third period and 14th group | C. Third period and 16th group | D. None of these

Answer: A

147. Which element has the highest melting point?

A. Na | B. K | C. Rb | D. Cs

Answer: A

148. 2nd and 3rd period called:

A. Short | B. Normal | C. Long | D. Very long

Answer: B

149. Number of elements in very long periods are:

A. 32 | B. 34 | C. 36 | D. 38

Answer: A

150. Transition element belongs to groups:

A. 7 to 8 | B. 3 to 12 | C. 12 to 18 | D. 7 to 18

Answer: B

151. Who discovered the atomic number?

A. Mendeleev | B. Moseley | C. Bohr | D. Dalton

Answer: B

152. What is trend in the reactivity of alkali metals?

A. Reactivity decrease in down the group | B. Reactivity increases down the group | C. Reactivity remains same | D. None of these

Answer: B

153. What is the nature of group-II metal oxide?

A. Acidic | B. Basic | C. Neutral | D. Amphoteric

Answer: B

154. Group of halogen family is:

A. 7 | B. 17 | C. 20 | D. 24

Answer: B

155. Identify the element that is larger in size among the following:

A. Li | B. Na | C. K | D. H

Answer: C

156. The tendency of an element to lose electron is called:

A. Electron affinity | B. Electronegativity | C. Electropositivity | D. Ionization energy

Answer: C

157. Which is the softest metal?

A. Na | B. Ca | C. Al | D. Zn

Answer: A

158. Outmost shell electronic configuration of Halogens is:

A. ns^2np^5 | B. ns^2np^2 | C. ns^2np^3 | D. ns^2np^4

Answer: A

159. Which oxide is the most basic oxide?

A. Na_2O | B. Li_2O | C. MgO | D. CO

Answer: A

160. Which group elements are the most reactive elements?

A. Transition metal group | B. First group | C. Second group | D. Third group

Answer: B

161. All the halogens exist as:

A. Monoatomic molecules | B. Diatomic molecules | C. Triatomic molecules | D. Polyatomic molecules

Answer: B

162. The gain of electron is called:

A. Oxidation | B. Reduction | C. Hydrolysis | D. Galvanizing

Answer: B

163. Correct order of oxidizing power of halogens is:

A. $Cl_2 > Br_2 > I_2$ | B. $Cl_2 > I_2 > Br_2$ | C. $I_2 > Cl_2 > Br_2$ | D. $Br_2 > Cl_2 > I_2$

Answer: A

164. Which one of the following is yellowish green gas?

A. F_2 | B. Cl_2 | C. Br_2 | D. I_2

Answer: B

165. The element present at the centre of the modern periodic table from group 3 to group 12 are called:

A. Transition elements | B. Halogens | C. Noble metals | D. Alkali metals

Answer: A

166. Which one of the following transition metal is used as catalyst for preparation of ammonia?

A. Ni | B. N_2O_5 | C. Fe | D. Cu

Answer: C

167. Hydrogenation of oil is carried out in the presence of:

A. Ni | B. Co | C. V_2O_5 | D. Fe

Answer: A

168. Outermost shell electronic configuration of noble gases is:

A. s^2p^6 | B. s^2p^5 | C. s^2p^4 | D. s^2p^3

Answer: A

169. Elements of group 18 are:

A. Noble metals | B. Noble non-metals | C. Alkali metals | D. None of these

Answer: B

170. Identify the noble gas that has 2 electrons in their valence shell:

A. He | B. Ne | C. Ar | D. Sr

Answer: A

171. All noble gases are:

A. Monoatomic | B. Diatomic | C. Triatomic | D. Polyatomic

Answer: A

172. The branch of chemistry which deals with the study of chemicals and other metal pollutants in the environment is called:

A. Environmental chemistry | B. Inorganic chemistry | C. Organic chemistry | D. Biochemistry

Answer: A

173. Blanket of air around the Earth is called:

A. Atmosphere | B. Hydrosphere | C. Lithosphere | D. Biosphere

Answer: A

174. Percentage of nitrogen in atmosphere is:

A. 21 | B. 78 | C. 0.04 | D. 15

Answer: B

175. Identify the major constituents of atmosphere:

A. N_2 , CO_2 | B. O_2 , CO_2 | C. N_2 , O_2 | D. He, O_2

Answer: C

176. The percentage of oxygen in atmosphere is:

A. 78% | B. 21% | C. 25% | D. 50%

Answer: B

177. The substances which cause air Pollution are called:

A. Destructive elements | B. Pollutants | C. Contaminants | D. Carbohydrates

Answer: B

178. Concentration of pollutants is expressed in:

A. pm | B. ppm | C. mm | D. nm

Answer: B

179. Indicate the source of sulphur which is responsible for the presence of oxides of sulphur in the atmosphere.

A. Waste gases from digestion of vegetation | B. Decomposition of animals | C. Combustion of fossil fuels | D. Photochemical smog

Answer: C

180. Why is smog not felt in summer?

A. Because fog is not present in summer | B. Because due to heat... smoke rises up | C. smoke and fog

cannot mix... | D. less fossil fuels are burnt...

Answer: A

181. Acid rain makes soil:

A. More acidic | B. Basic | C. Neutral | D. None of these

Answer: A

182. Which of the following is a greenhouse gas?

A. O_2 | B. N_2 | C. CO_2 | D. Cl_2

Answer: C

183. Chemical used to neutralized acidic soil is:

A. Soda lime | B. Soda ash | C. Caustic Soda | D. Washing Soda

Answer: A

184. In which area there is a greater possibility of acid rain?

A. Around village | B. Around big cities | C. Around industrial area | D. Around waterbodies

Answer: C

185. Which gases contribute towards the formation of acid rain?

A. Oxides of carbon | B. Oxides of sulphur | C. Oxides of nitrogen | D. Both the oxides of nitrogen and sulphur

Answer: D

186. Which of the following is acid rain?

A. NO_x | B. SO_x | C. CO_2 | D. All of these

Answer: B

187. The essential element present in organic compounds is:

A. Oxygen | B. Nitrogen | C. Carbon | D. Hydrogen

Answer: C

188. Which one of the following is saturated compound?

A. C_2H_6 | B. CH_3Cl | C. CH_3OH | D. NH_3

Answer: A

189. The simplest alkane is:

A. Methane | B. Ethane | C. Propane | D. Butane

Answer: A

190. Which other atom is almost always present along with carbon atom in all organic compounds?

A. Oxygen | B. Nitrogen | C. Hydrogen | D. Halogen

Answer: C

191. Identify the parent hydrocarbon:

A. C_2H_2 | B. C_2H_4 | C. CH_4 | D. C_2H_6

Answer: C

192. Which part of IUPAC system tells us about class of organic compound?

A. Root | B. Suffix | C. Prefix | D. All of these

Answer: B

193. According to Bronsted-Lowry, an example of Base is:

A. OH^- | B. CN^- | C. Cl^- | D. All of these

Answer: D

194. Which one of the following does not react with chemical reagents like $KMnO_4$?

A. Alkane | B. Alkene | C. Alkyne | D. None of these

Answer: A

195. Elements of group 18 are:

A. Noble gases | B. Actinides | C. Alkali metals | D. None of these

Answer: A

196. The catalyst used for reduction of alkene and alkyne is:

A. Fe | B. P | C. Ni | D. C

Answer: C

197. The word paraffin's mean:

A. High affinity | B. Little affinity | C. Moderate affinity | D. None of these

Answer: B

198. Alkanes react with excess oxygen to produce:

A. CO_2 and H_2 | B. CO_2 and H_2O | C. CO and H_2O | D. CO and H_2

Answer: B

199. Replacement of hydrogen of Alkane by an atom or a group of atoms are called:

A. Addition reaction | B. Substitution reaction | C. Sublimation | D. Decomposition

Answer: B

200. Which product will be formed when ethyl bromide ($\text{CH}_3\text{CH}_2\text{Br}$) is treated with Zn/HCl ?

A. CH_4 | B. $\text{CH}_3\text{-CH}_3$ | C. $\text{CH}_3\text{-CH}_2\text{-CH-CH}_3$ | D. $\text{CH}_3\text{-CH}_2\text{-CH}_3$

Answer: B

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